Final Answer key of Engineering Entrance Competitive Examination (Lateral Entry) – 2026 held on 28.06.2026

Set Code	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18	Q19	Q20	Q21	Q22	Q23	Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40	Q41	Q42	Q43	Q44	Q45	Q46	Q47	Q48	Q49	Q50
A	A	B	B	B	A	C	A	A	C	B	B	B	B	B	C	B	B	A	A	B	B	B	C	A	A	D	Deleted	B	C	A	C	A	B	B	C	C	B	A	D	C	A	B	A	D	B	B	B	B	C	A
	Q51	Q52	Q53	Q54	Q55	Q56	Q57	Q58	Q59	Q60	Q61	Q62	Q63	Q64	Q65	Q66	Q67	Q68	Q69	Q70	Q71	Q72	Q73	Q74	Q75	Q76	Q77	Q78	Q79	Q80	Q81	Q82	Q83	Q84	Q85	Q86	Q87	Q88	Q89	Q90	Q91	Q92	Q93	Q94	Q95	Q96	Q97	Q98	Q99	Q100
	C	A	A	C	C	C	A	C	C	B	B	D	B	B	B	A	B	A	D	C	A	B	B	B	A & B	B	B	B	B	C	B	A	A	B	C	B	B	D	B	B	B	D	D	B	B	B	D	D	D	B
	Q101	Q102	Q103	Q104	Q105	Q106	Q107	Q108	Q109	Q110	Q111	Q112	Q113	Q114	Q115	Q116	Q117	Q118	Q119	Q120	Q121	Q122	Q123	Q124	Q125	Q126	Q127	Q128	Q129	Q130	Q131	Q132	Q133	Q134	Q135	Q136	Q137	Q138	Q139	Q140	Q141	Q142	Q143	Q144	Q145	Q146	Q147	Q148	Q149	Q150
	D	A	C	C	D	C	B	A	D	D	D	C	D	B	A	B	B	C	B	B	D	B	A	A	C	A	A	B	C	B	A	D	A	C	A	B	D	B	B	B	A	D	B	A	C	D

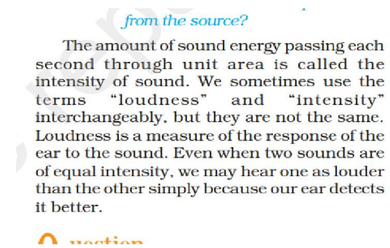
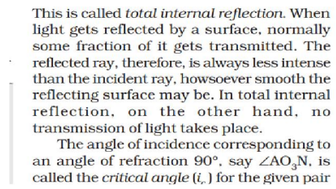
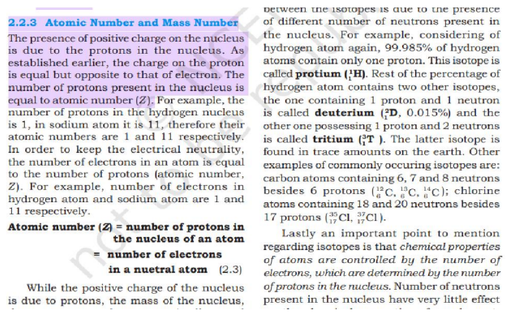
Set Code	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18	Q19	Q20	Q21	Q22	Q23	Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40	Q41	Q42	Q43	Q44	Q45	Q46	Q47	Q48	Q49	Q50
B	C	B	B	A	B	B	A	B	C	B	A	D	Deleted	A	C	A	B	A	C	B	C	B	B	C	D	A	A	B	C	D	B	A	B	B	B	A	C	B	A	B	B	C	A	A	A	B	C	B	B	B
	Q51	Q52	Q53	Q54	Q55	Q56	Q57	Q58	Q59	Q60	Q61	Q62	Q63	Q64	Q65	Q66	Q67	Q68	Q69	Q70	Q71	Q72	Q73	Q74	Q75	Q76	Q77	Q78	Q79	Q80	Q81	Q82	Q83	Q84	Q85	Q86	Q87	Q88	Q89	Q90	Q91	Q92	Q93	Q94	Q95	Q96	Q97	Q98	Q99	Q100
	B	A	D	A	A	B	C	B	B	B	B	A & B	B	C	B	A	B	B	A	B	C	B	B	D	B	D	B	B	D	B	B	D	D	B	D	A	A	C	C	C	C	C	C	A	B	B	B	D	B	B
	Q101	Q102	Q103	Q104	Q105	Q106	Q107	Q108	Q109	Q110	Q111	Q112	Q113	Q114	Q115	Q116	Q117	Q118	Q119	Q120	Q121	Q122	Q123	Q124	Q125	Q126	Q127	Q128	Q129	Q130	Q131	Q132	Q133	Q134	Q135	Q136	Q137	Q138	Q139	Q140	Q141	Q142	Q143	Q144	Q145	Q146	Q147	Q148	Q149	Q150
	A	B	B	B	B	B	D	C	A	A	B	A	A	C	C	B	A	B	A	C	D	D	A	C	A	B	D	B	A	B	A	D	B	A	C	D	B	A	C	D	D	C	B	C	D	D	A	C	D	D

Set Code	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18	Q19	Q20	Q21	Q22	Q23	Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40	Q41	Q42	Q43	Q44	Q45	Q46	Q47	Q48	Q49	Q50
C	A	B	C	C	A	C	B	B	C	B	A	D	B	C	A	B	A	D	B	B	B	C	A	A	B	B	B	A	C	A	A	C	B	B	B	B	B	A	D	C	A	B	A	D	B	B	B	D	A	A
	Q51	Q52	Q53	Q54	Q55	Q56	Q57	Q58	Q59	Q60	Q61	Q62	Q63	Q64	Q65	Q66	Q67	Q68	Q69	Q70	Q71	Q72	Q73	Q74	Q75	Q76	Q77	Q78	Q79	Q80	Q81	Q82	Q83	Q84	Q85	Q86	Q87	Q88	Q89	Q90	Q91	Q92	Q93	Q94	Q95	Q96	Q97	Q98	Q99	Q100
	A	B	C	B	B	D	B	B	B	D	D	B	B	B	D	D	D	B	A	C	A	C	C	C	C	A	C	B	B	D	B	B	B	A	D	A	B	B	C	A	B	B	B	A & B	B	C	B	B	B	A
	Q101	Q102	Q103	Q104	Q105	Q106	Q107	Q108	Q109	Q110	Q111	Q112	Q113	Q114	Q115	Q116	Q117	Q118	Q119	Q120	Q121	Q122	Q123	Q124	Q125	Q126	Q127	Q128	Q129	Q130	Q131	Q132	Q133	Q134	Q135	Q136	Q137	Q138	Q139	Q140	Q141	Q142	Q143	Q144	Q145	Q146	Q147	Q148	Q149	Q150
	B	A	B	C	C	D	A	A	C	A	D	D	B	A	B	A	D	B	B	C	D	B	A	C	D	A	C	B	C	D	D	A	D	D	D	C	B	B	B	A	B	D	C	B	A	B	A	A	C	A

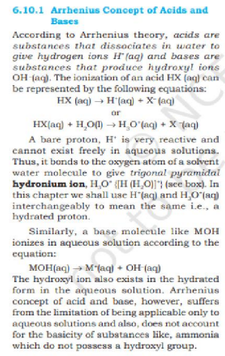
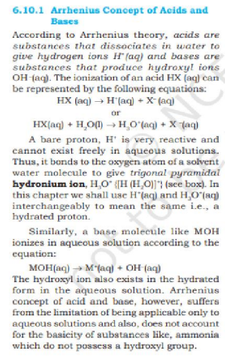
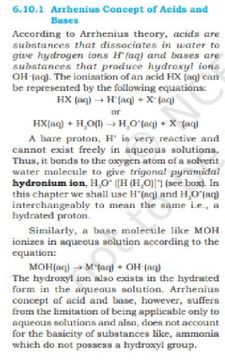
Set Code	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8	Q9	Q10	Q11	Q12	Q13	Q14	Q15	Q16	Q17	Q18	Q19	Q20	Q21	Q22	Q23	Q24	Q25	Q26	Q27	Q28	Q29	Q30	Q31	Q32	Q33	Q34	Q35	Q36	Q37	Q38	Q39	Q40	Q41	Q42	Q43	Q44	Q45	Q46	Q47	Q48	Q49	Q50
D	D	A	C	A	B	A	D	B	B	B	B	A	C	B	A	B	B	C	A	A	A	B	C	B	B	B	B	C	B	A	B	B	A	B	C	B	Deleted	A	A	D	B	C	A	A	C	B	B	C	B	C
	Q51	Q52	Q53	Q54	Q55	Q56	Q57	Q58	Q59	Q60	Q61	Q62	Q63	Q64	Q65	Q66	Q67	Q68	Q69	Q70	Q71	Q72	Q73	Q74	Q75	Q76	Q77	Q78	Q79	Q80	Q81	Q82	Q83	Q84	Q85	Q86	Q87	Q88	Q89	Q90	Q91	Q92	Q93	Q94	Q95	Q96	Q97	Q98	Q99	Q100
	B	D	B	B	D	B	B	D	D	B	D	C	A	A	C	C	C	A	C	C	B	B	B	D	B	B	D	A	B	A	C	A	B	B	B	A & B	B	B	B	B	C	A	B	B	A	B	C	B	B	D
	Q101	Q102	Q103	Q104	Q105	Q106	Q107	Q108	Q109	Q110	Q111	Q112	Q113	Q114	Q115	Q116	Q117	Q118	Q119	Q120	Q121	Q122	Q123	Q124	Q125	Q126	Q127	Q128	Q129	Q130	Q131	Q132	Q133	Q134	Q135	Q136	Q137	Q138	Q139	Q140	Q141	Q142	Q143	Q144	Q145	Q146	Q147	Q148	Q149	Q150
	C	A	D	A	B	A	B	D	D	B	B	A	D	B	A	C	D	A	C	B	C	D	C	A	D	D	D	C	D	B	B	A	B	D	C	B	B	B	A	A	C	A	A	A	B	C	B	D	A	C

Note :- 1) Set-A-Q27 / Set-B-Q13 / Set-C-Q48 / Set-D-Q37: Marks will be awarded to all candidates for these questions.
2) Set-A-Q75 / Set-B-Q62 / Set-C-Q94 / Set-D-Q86: Both Answer A and Answer B are correct. Therefore, candidates who selected either Answer A or Answer B will be awarded marks.

Clarification of Objections

Section Number	Section Name	Master Set Q.No.	Bklet/ Q. No.	Question Description as per Master Set (Set-A)	Answer key	Objection	Remarks Given by Candidates	Objection Status Valid/ Invalid	Final Clarification and Reference Given by Subject Expert	Final Answer Given By Subject Expert
1	Physics	27	D-37	**27.** The amount of sound energy passing each sound through unit area is called _____. (A) density of sound (B) velocity of sound (C) intensity of sound (D) speed of sound इकाई क्षेत्रफल से गुज़रने वाली ध्वनि ऊर्जा की मात्रा को _____ कहा जाता है । (A) ध्वनि का घनत्व (B) ध्वनि का वेग (C) ध्वनि की तीव्रता (D) ध्वनि की गति	C	Incorrect Question	Printing/Typographical error in the question statement Elaboration: The phrase "passing each sound through unit area" is not meaningful in physics terminology. The standard definition of sound intensity is "the amount of sound energy passing through unit area per second". It appears that the word "second" was mistakenly replaced by "sound". Due to this printing error, the wording of the question is defective and may reasonably confuse candidates regarding the intended physical quantity. Thus, I respectfully request the Board to treat this as a defective question and take appropriate action, including awarding appropriate marks to affected candidates, if deemed fit under the Board's examination policy.	Valid	The definition of sound intensity would be the amount of sound energy per second passing through a unit surface area normal to the direction of propagation. It seems that the word “sound” should have been “second.” Std 9, Science, Ch 12 Sound, Page 166  *from the source?* The amount of sound energy passing each second through unit area is called the intensity of sound. We sometimes use the terms “loudness” and “intensity” interchangeably, but they are not the same. Loudness is a measure of the response of the ear to the sound. Even when two sounds are of equal intensity, we may hear one as louder than the other simply because our ear detects it better.	Question Should be Deleted
1	Physics	33	D-46	**33.** When a ray of light is reflected by a surface, the intensity of reflected ray (A) always increases. (B) always decreases. (C) remains unchanged. (D) either increases or decreases. जब प्रकाश की किरण किसी सतह से परावर्तित होती है, तो परावर्तित किरण की तीव्रता (A) हमेशा बढ़ती है । (B) हमेशा घटती है । (C) अपरिवर्तित रहती है । (D) या तो बढ़ती है या घटती है ।	B	Incorrect Question	Ambiguous question due to omission of the nature of the reflecting surface	Invalid	In reflection from a real surface, the intensity of the reflected ray is less than the incident ray because a part of the incident light is absorbed, scattered, or transmitted by the surface. Therefore, the intensity of the reflected ray decreases. In standard physics, unless the word "ideal" or "perfectly reflecting" is explicitly mentioned, you must assume the laws of practical, real-world physics. Std 12, Physics, Ch 9, Ray Optics and Optical Instruments, Page 320  This is called *total internal reflection*. When light gets reflected by a surface, normally some fraction of it gets transmitted. The reflected ray, therefore, is always less intense than the incident ray, howsoever smooth the reflecting surface may be. In total internal reflection, on the other hand, no transmission of light takes place. The angle of incidence corresponding to an angle of refraction 90°, say $\angle AO_3N$, is called the *critical angle* (i_c) for the given pair	(B)
2	Chemistry	53	D-64	**53.** The introduction of neutron in nucleus of an atom would lead to the change in (A) Atomic weight (B) Atomic number (C) Number of electrons around a nucleus (D) None of these किसी परमाणु के नाभिक में न्यूट्रॉन डालने से किसमें बदलाव आएगा ? (A) परमाणु भार (B) परमाणु क्रमांक (C) नाभिक के चारों ओर इलेक्ट्रॉनों की संख्या (D) इनमें से कोई नहीं	A	Mass Number ≠ Atomic Mass ≠ Atomic Weight	Invalid	The mass of the nucleus determined due to protons and neutrons only. Protons and neutrons present in the nucleus are collectively known as nucleons which decides the Atomic weight/ mass number. When a neutron is added to the nucleus of an atom, the number of protons remains unchanged; therefore, the atomic number does not change. The number of electrons around the nucleus also does not necessarily change. However, since a neutron has mass, addition of a neutron increases the mass of the atom. Hence, the atomic weight/atomic mass changes.  **2.2.3 Atomic Number and Mass Number** The presence of positive charge on the nucleus is due to the protons in the nucleus. As established earlier, the charge on the proton is equal but opposite to that of electron. The number of protons present in the nucleus is equal to atomic number (*Z*). For example, the number of protons in the hydrogen nucleus is 1, in sodium atom it is 11, therefore their atomic numbers are 1 and 11 respectively. In order to keep the electrical neutrality, the number of electrons in an atom is equal to the number of protons (atomic number, Z). For example, number of electrons in hydrogen atom and sodium atom are 1 and 11 respectively. **Atomic number (Z) = number of protons in the nucleus of an atom = number of electrons in a nuetral atom** (2.3) While the positive charge of the nucleus is due to protons, the mass of the nucleus, between the isotopes is due to the presence of different number of neutrons present in the nucleus. For example, considering of hydrogen atom again, 99.985% of hydrogen atoms contain only one proton. This isotope is called **protium** ($_1^1H$). Rest of the percentage of hydrogen atom contains two other isotopes, the one containing 1 proton and 1 neutron is called **deuterium** ($_1^2D$, 0.015%) and the other one possessing 1 proton and 2 neutrons is called **tritium** ($_1^3T$). The latter isotope is found in trace amounts on the earth. Other examples of commonly occuring isotopes are: carbon atoms containing 6, 7 and 8 neutrons besides 6 protons ($_6^{12}C$, $_6^{13}C$, $_6^{14}C$); chlorine atoms containing 18 and 20 neutrons besides 17 protons ($_{17}^{35}Cl$, $_{17}^{37}Cl$). Lastly an important point to mention regarding isotopes is that *chemical properties of atoms are controlled by the number of electrons, which are determined by the number of protons in the nucleus.* Number of neutrons present in the nucleus have very little effect NCERT Book Class-XI Chemistry Part-1, Chapter: 02, Atomic Structure, Page No.352.2.3 Atomic Number and Mass Number	(A)

Section Number	Section Name	Master Set Q.No.	Bklet/ Q. No.	Question Description as per Master Set (Set-A)	Answer key	Objection	Remarks Given by Candidates	Objection Status Valid/ Invalid	Final Clarification and Reference Given by Subject Expert	Final Answer Given By Subject Expert
2	Chemistry	54	D-65	**54.** The number of electrons, protons and neutrons in a species are equal to 18, 16 and 16 respectively. What is the proper symbol of that species ? एक स्पीशीज़ में इलेक्ट्रॉनों, प्रोटॉनों और न्यूट्रॉनों की संख्या क्रमशः 18, 16 और 16 है । उस स्पीशीज़ का सही प्रतीक क्या है ? (A) Al (B) Al^{+3} (C) S^{-2} (D) S	C		Oxidation states (Al^{+3}, S^{-2}) & Charged Ion Species (Al^{3+}, S^{2}) are not the same thing.	Invalid	The question asks for the proper symbol of the species on the basis of the given numbers of electrons, protons and neutrons. Since the species has 16 protons, it is sulphur. Since it has 18 electrons, it has two electrons more than the neutral sulphur atom and therefore carries a charge of 2–. Hence the correct symbol is S^{2-}. This question is about identifying the charged ionic species, not assigning oxidation state. Therefore, the objection regarding oxidation state notation is not applicable. **Solution** In this case, $^{80}_{35}Br$, Z = 35, A = 80, species is neutral Number of protons = number of electrons = Z = 35 Number of neutrons = 80 – 35 = 45, (equation 2.4) **Problem 2.2** The number of electrons, protons and neutrons in a species are equal to 18, 16 and 16 respectively. Assign the proper symbol to the species. **Solution** The atomic number is equal to number of protons = 16. The element is sulphur (S). Atomic mass number = number of protons + number of neutrons = 16 + 16 = 32 Species is not neutral as the number of protons is not equal to electrons. It is anion (negatively charged) with charge equal to excess electrons = 18 – 16 = 2. Symbol is $^{32}_{16}S^{2-}$. **Note :** Before using the notation $^{A}_{Z}X$, NCERT Book Class-XI Chemistry Part-1, Chapter: 02, Intext Question, Problem 2.2 "The number of electrons, protons and neutrons in a species are equal to 18, 16 and 16 respectively. Assign the proper symbol to the species" NCERT Book Class-XI Chemistry Part-1, Chapter: 02, Atomic Structure, Problem 2.2 (Intext Question) Page No.36	(c)
2	Chemistry	62	A-62	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Respected Sir/Madam, I respectfully submit this objection regarding Question No. 62 of Set A in the provisional answer key. The provisional answer key marks Option (D) "Ionization in aqueous solution" as the correct answer. However, according to the Arrhenius acid-base theory, the strength of an acid or base depends on the extent (degree) of dissociation in aqueous solution. Strong acids and bases dissociate almost completely in water, while weak acids and bases dissociate only partially. Therefore, the strength is determined by the extent of dissociation. Option (A), "Extent of dissociation in aqueous solution," directly expresses this concept and is widely used in chemistry textbooks. In view of the above, I kindly request the Board to review Question No. 62 and consider Option (A) as the correct answer, or take appropriate action if the question is found to be ambiguous. Thank you for your time and consideration.	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ ([H $(H_2O)]^+$] (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)
2	Chemistry	62	C-80	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) and (D) is correct answer	In Set C, question number 80 the answer would be both A & D	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ ([H $(H_2O)]^+$] (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)

Section Number	Section Name	Master Set Q.No.	Bklet/ Q. No.	Question Description as per Master Set (Set-A)	Answer key	Objection	Remarks Given by Candidates	Objection Status Valid/ Invalid	Final Clarification and Reference Given by Subject Expert	Final Answer Given By Subject Expert
2	Chemistry	62	A-62	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Good morning sir Set A question no. 62 has been answered incorrectly, it has been given in the answer key but ar swer a hoga	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 	(D)
2	Chemistry	62	A-62	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Respected Sir/Madam, I respectfully submit this objection regarding Question No. 62 of Set A in the Provisional Answer Key of D2D 2026. Your provisional answer key has marked Option (D) - "Ionization in aqueous solution" as the correct answer. But, I found that this is incorrect in many sources. Because According to the Arrhenius acid-base theory, the strength of an acid or base depends on the extent (degree) of dissociation in aqueous solution. Strong acids and bases dissociate almost completely in water, whereas weak acids and bases dissociate only partially. Therefore, the strength of an acid or base is determined by the extent of dissociation, not merely by the process of ionization. Option (A) - "Extent of dissociation in aqueous solution" directly expresses this concept and is the standard terminology used in chemistry textbooks. In view of the above, I kindly request the Board to review Question No. 62 and consider Option (A) as the correct answer, or take appropriate action if the question is found to be ambiguous. Thank you for your time and consideration.	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 	(D)
2	Chemistry	62	C-80	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Respected sir/Madam, I respectfully submitted this objection regarding Question NO:80 of set C in provisional answer key of D2D 2026. Your provisional answer key has marked option(D)_"ionization in aqueous solution" as correct answer. But, I found that this is incorrect in according to your question. Arrhenius Acid-Base Theory Based on the Arrhenius theory, the correct answer to question 80 is: (A) extent of dissociation in aqueous solution Explanation The Arrhenius theory defines acids as substances that dissociate in water to yield hydrogen ions (\(H^{+}\)) and bases as substances that dissociate in water to yield hydroxide ions (\(OH^{-}\)). Strong acids/bases dissociate or ionize completely in aqueous solution. Weak acids/bases dissociate only partially. Therefore, the "strength" is directly measured by the extent of dissociation (how much of the substance breaks apart into ions) when dissolved in water.	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 	(D)

Section Number	Section Name	Master Set Q.No.	Bklet/ Q. No.	Question Description as per Master Set (Set-A)	Answer key	Objection	Remarks Given by Candidates	Objection Status Valid/ Invalid	Final Clarification and Reference Given by Subject Expert	Final Answer Given By Subject Expert
2	Chemistry	62	D-74	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Respected Sir/Madam, I am a candidate who appeared for the Engineering Entrance Competitive Examination (Lateral Entry) 2026, and I wish to raise an objection against the provisional answer key published for the exam. Candidate details: Roll No.: 2602104568 Set: D Question No.: 74 Correct answer - (option A) Given (option D), which is wrong Question: "According to Arrhenius acid-base theory, the strength of acid and base depends on the -" (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton (C) magnitude of donating proton (D) ionization in aqueous solution	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX\,(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ ($[H(H_2O)]^+$) (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)
2	Chemistry	62	B-98	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	I would like to bring to your kind attention that in the official answer key for Set B published by the correct answer for question no. 98 has been marked as Option D. However, based on the attached supporting evidence, Option A appears to be the more appropriate and correct answer. I have attached the relevant proof for your kind reference. I kindly request you to review the same and take the necessary action. Kind regards, Aniket Kumar Singh Document from Aniket Singh.pdf 944K	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX\,(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ ($[H(H_2O)]^+$) (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)
2	Chemistry	62	C-80	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	(A) is correct answer	Qno. 80. Correct answer: A. extent of dissociation in aqueous solution* "Explanation with logic:* "Step 1: Arrhenius definition* Arrhenius 1884 theory only works for water: 1. "Acid gives H^+ in water: $HA \rightarrow H^+ + A^-$ 2. "Base* gives $OH^-S in water: SBOH \rightarrow B^+ + OH^-$ "Step 2: How does he define "strength"?* Strength = power to give H^+ or OH^- ions. More H^+ ions = Stronger acid More OH^- ions = Stronger base "Step 3: What controls number of ions?* The number of ions depends on "extent of dissociation" = how much % of acid/base molecules break into ions in water. Formula: $\alpha = \frac{\text{No. of molecules dissociated}}{\text{Total no. of	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190 **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX\,(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ ($[H(H_2O)]^+$) (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)

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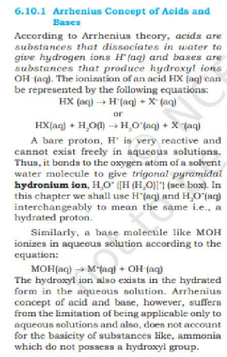
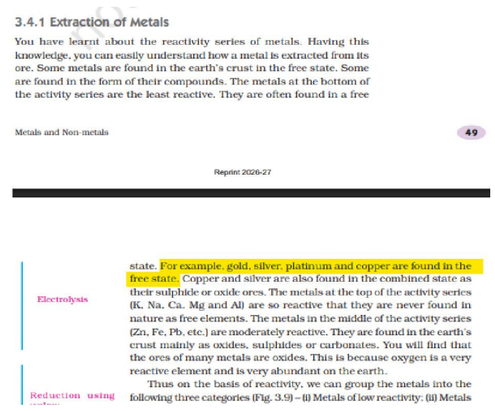
2	Chemistry	62	A-62	**62.** According to Arrhenius acid–base theory, the strength of acid and base depends on the (A) extent of dissociation in aqueous solution (B) magnitude of accepting proton. (C) magnitude of donating proton. (D) ionization in aqueous solution. आरेनिअस एसिड–बेस थ्योरी (अम्ल–क्षार सिद्धांत) के अनुसार, अम्ल और क्षार की ताकत किस पर निर्भर करती है ? (A) जलीय विलयन में वियोजन की मात्रा (B) प्रोटॉन स्वीकार करने की क्षमता । (C) प्रोटॉन दान करने की क्षमता । (D) जलीय विलयन में आयनीकरण ।	D	Incorrect Key	Ambiguous chemical nomenclature in option (d)	Invalid	As per Arrhenius theory, acids and bases are defined on the basis of their ionization in aqueous solution. Acids ionize to give hydrogen/hydronium ions and bases ionize to give hydroxyl ions. According to Arrhenius acid-base theory, acids and bases produce ions in aqueous solution. Acids ionize in water to give hydrogen ions/hydronium ions, while bases ionize in water to give hydroxyl ions. The reference of NCERT also clearly explains the ionization of acid HX and base MOH in aqueous solution. Therefore, in the context of Arrhenius theory, the strength of acids and bases is related to their ionization in aqueous solution. Ref: Std 11, Chemistry, Unit 6, Equilibrium, Page 190  **6.10.1 Arrhenius Concept of Acids and Bases** According to Arrhenius theory, *acids are substances that dissociates in water to give hydrogen ions $H^+(aq)$ and bases are substances that produce hydroxyl ions* $OH^-(aq)$. The ionization of an acid HX (aq) can be represented by the following equations: $HX\,(aq) \rightarrow H^+(aq) + X^-(aq)$ or $HX(aq) + H_2O(l) \rightarrow H_3O^+(aq) + X^-(aq)$ A bare proton, H^+ is very reactive and cannot exist freely in aqueous solutions. Thus, it bonds to the oxygen atom of a solvent water molecule to give *trigonal pyramidal* **hydronium ion**, H_3O^+ [$(H(H_2O)^+)$] (see box). In this chapter we shall use $H^+(aq)$ and $H_3O^+(aq)$ interchangeably to mean the same i.e., a hydrated proton. Similarly, a base molecule like MOH ionizes in aqueous solution according to the equation: $MOH(aq) \rightarrow M^+(aq) + OH^-(aq)$ The hydroxyl ion also exists in the hydrated form in the aqueous solution. Arrhenius concept of acid and base, however, suffers from the limitation of being applicable only to aqueous solutions and also, does not account for the basicity of substances like, ammonia which do not possess a hydroxyl group.	(D)
2	Chemistry	75	D-86	**75.** Among the following select the metal found free in nature नीचे दी गई धातुओं में से उस धातु को चुनिए जो प्रकृति में मुक्त अवस्था में पाई जाती है : (A) Au (B) Cu (C) Na (D) Hg	A	More than one correct answer	More than one scientifically correct option	Valid	A few elements like carbon, sulphur, gold and noble gases, occur in free state while others are found in combined forms in the earth's crust. NCERT Std 10, Science, Ch 3,Page 49-50  **3.4.1 Extraction of Metals** You have learnt about the reactivity series of metals. Having this knowledge, you can easily understand how a metal is extracted from its ore. Some metals are found in the earth's crust in the free state. Some are found in the form of their compounds. The metals at the bottom of the activity series are the least reactive. They are often found in a free Metals and Non-metals 49 Reprint 2026-27 Electrolysis Reduction using state. For example, gold, silver, platinum and copper are found in the free state. Copper and silver are also found in the combined state as their sulphide or oxide ores. The metals at the top of the activity series (K, Na, Ca, Mg and Al) are so reactive that they are never found in nature as free elements. The metals in the middle of the activity series (Zn, Fe, Pb, etc.) are moderately reactive. They are found in the earth's crust mainly as oxides, sulphides or carbonates. You will find that the ores of many metals are oxides. This is because oxygen is a very reactive element and is very abundant on the earth. Thus on the basis of reactivity, we can group the metals into the following three categories (Fig. 3.9) – (i) Metals of low reactivity; (ii) Metals	(A) & (B)
2	Chemistry	94	A-94	**94.** How much Sulphur is added during vulcanization to get rubber for tyre ? (A) 30 Percent (B) 5 Percent (C) 55 Percent (D) 3 Percent टायर के लिए रबर बनाने की वल्कनीकरण प्रक्रिया में कितना सल्फर मिलाया जाता है ? (A) 30 प्रतिशत (B) 5 प्रतिशत (C) 55 प्रतिशत (D) 3 प्रतिशत	B	(B) and (D) is correct answer	Dear sir/mam I would like to draw your kind attention towards the answer key published for Engineering lateral entry examination 2026. In this I found a question whose answer is not correct. So I'm mentioning the question along with Set, alongwith I am also sharing the question with the proof of correct answer and request you to review it by your own side to the depth. SET-A Question number - 94 In this question it is asked that what percentage of sulphur is added during vulcanisation to get rubber for Tyre? Options given, A-30% B-5% C-55% D-3% It's answer as given in answer key is 5% Objection - This answer (5%) is not perfect because a tyre should have sulphur addition between (3-5)%. In the options provided both 3 & 5% is in option B and D so both the answers should be correct. I request the concerned authority to take an eye on this by reviewing it from renowned books of chemistry (Nirali, Jhunjhunwala) in every single book 3-5% is mentioned or by the help of Google, as a single marks can decide the future of a child. Yours hopefully	Invalid	Although Vulcanisation of rubber is heationg it with different concentration of sulphur but for manufacturing TYRE RUBBER, 5% of sulphur is used. It is clearly mentioned tn NCERT Text book. https://cdn1.byjus.com/wp-content/uploads/2019/10/NCERT-Book-for-Class-12-Chemistry-Chapter-15.pdf OLD NCERT Book Class-XII Chemistry Part-II, Chapter: 15, POLYMERS, Page No.441 TOPIC: 15.2.4 : Rubber, Vulcanisation of rubber at a temperature range between 373 K to 415 K. On vulcanisation, sulphur forms cross links at the reactive sites of double bonds and thus the rubber gets stiffened. In the manufacture of tyre rubber, 5% of sulphur is used as a crosslinking agent. The probable structures of vulcanised rubber molecules are depicted below:	(B)

Section Number	Section Name	Master Set Q.No.	Bklet/ Q. No.	Question Description as per Master Set (Set-A)	Answer key	Objection	Remarks Given by Candidates	Objection Status Valid/ Invalid	Final Clarification and Reference Given by Subject Expert	Final Answer Given By Subject Expert
2	Chemistry	98	C-66	**98.** Which one of the following is not a greenhouse gas ? (A) Carbon dioxide (B) Ozone (C) Methane (D) Nitrogen oxide इनमें से कौन–सी ग्रीनहाउस गैस नहीं है ? (A) कार्बन डाइऑक्साइड (B) ओज़ोन (C) मीथेन (D) नाइट्रोजन ऑक्साइड	D	Incorrect Question	Qno. 66, Which one of the following is not a greenhouse gas?* *Correct answer: Trick question - All of them are greenhouse gases "Explanation: Greenhouse gases = gases that trap heat in atmosphere by absorbing infrared radiation. Check each option: 1. *(A) Carbon dioxide CO_2* - Yes, main greenhouse gas.76% of total GHG emissions 2. *(B) Ozone SO_3$ - Yes. Ground-level ozone/tropospheric ozone is a greenhouse gas. Only stratospheric ozone "good ozone" protects from UV. But chemically O_3 does trap heat 3. (C) Methane CH_4* - Yes. 2nd biggest after CO_2. Traps 25x more heat than CO_2 per molecule 4. "(D) Nitrogen oxide - Yes. N_20 = Nitrous oxide is a greenhouse gas. NO_x also contribute indirectly "Why confusion happens:* Books often teach "main GHGs = $CO_2, CH_4, N_20, CFCs$". Students then think ozone isn't one. But IPCC also lists SO_35 as greenhouse gas. So if your exam forces you to choose one, the question is wrong. If it's from a source that expects "Ozone", they mean "ozone layer isn't the main one we talk about for global warming". But scientifically, option B is still a greenhouse gas. Bottom line: By strict science, "none of the above" is correct.	Invalid	Carbon dioxide, ozone and methane are recognized greenhouse gases. The option “nitrogen oxide” in the question refers to common nitrogen oxides such as nitric oxide/nitrogen monoxide (NO) and nitrogen dioxide (NO_2), which are generally treated as air pollutants and not listed as principal greenhouse gases at this level. Nitrous oxide, N_2O, is a greenhouse gas; however, it is specifically called nitrous oxide, not simply “nitrogen oxide.” Since the question uses the term “nitrogen oxide” and not “nitrous oxide,” the intended answer is clear in the given context. • **Greenhouse gases(GHGs)** are a group of compounds that are responsible for **absorbing heat (longwave radiation) in the atmosphere, keeping the Earth's surface warmer by allowing direct sunlight (relative short wave energy) to reach the Earth's surface.** • **Most Prominent Green House Gases**: ◦ Water vapour ◦ Carbon Dioxide, CO2 ◦ Methane, CH4 ◦ **Nitrous Oxide, N_2O (and NOT Nitrogen Oxide)** ◦ Ozone, O3 • Other GHGs are **carbon monoxide, fluorinated gases, chlorofluorocarbons (CFCs), black carbon (soot), and brown carbon.** • Amongst all the **greenhouse gases, only water vapour can absorb both incoming (UV) and outgoing (infrared) radiation.** • If GHG emissions remain unchecked, by the end of this century, the global temperatures may increase by **5^{o}C, causing global warming.**	(D)
2	Chemistry	98	C-66	**98.** Which one of the following is not a greenhouse gas ? (A) Carbon dioxide (B) Ozone (C) Methane (D) Nitrogen oxide इनमें से कौन–सी ग्रीनहाउस गैस नहीं है ? (A) कार्बन डाइऑक्साइड (B) ओज़ोन (C) मीथेन (D) नाइट्रोजन ऑक्साइड	D	(B) is correct answer	Please find the attached google doc containing the consolidated list of objections (with supporting evidence) against the released Answer Keys of Engineering Entrance Competitive Examination (Lateral Entry) - 2026 held on 28.06.2026 (keys released on 01.07.2026) The objections are based on the question paper SET D, and includes question numbers 37, 46, 58, 74 & 86. JD2D 2026 - Answer Key Objections I have also attached below a pdf version of the google doc if you would rather use that instead. PDF JD2D 2026 - Answer Key Objections.pdf Regards, Priyanshu Kumar Exam Roll No.: 2602102188	Invalid	Carbon dioxide, ozone and methane are recognized greenhouse gases. The option “nitrogen oxide” in the question refers to common nitrogen oxides such as nitric oxide/nitrogen monoxide (NO) and nitrogen dioxide (NO_2), which are generally treated as air pollutants and not listed as principal greenhouse gases at this level. Nitrous oxide, N_2O, is a greenhouse gas; however, it is specifically called nitrous oxide, not simply “nitrogen oxide.” Since the question uses the term “nitrogen oxide” and not “nitrous oxide,” the intended answer is clear in the given context. • **Greenhouse gases(GHGs)** are a group of compounds that are responsible for **absorbing heat (longwave radiation) in the atmosphere, keeping the Earth's surface warmer by allowing direct sunlight (relative short wave energy) to reach the Earth's surface.** • **Most Prominent Green House Gases**: ◦ Water vapour ◦ Carbon Dioxide, CO2 ◦ Methane, CH4 ◦ **Nitrous Oxide, N_2O (and NOT Nitrogen Oxide)** ◦ Ozone, O3 • Other GHGs are **carbon monoxide, fluorinated gases, chlorofluorocarbons (CFCs), black carbon (soot), and brown carbon.** • Amongst all the **greenhouse gases, only water vapour can absorb both incoming (UV) and outgoing (infrared) radiation.** • If GHG emissions remain unchecked, by the end of this century, the global temperatures may increase by **5^{o}C, causing global warming.**	(D)
2	Chemistry	98	D-58	**98.** Which one of the following is not a greenhouse gas ? (A) Carbon dioxide (B) Ozone (C) Methane (D) Nitrogen oxide इनमें से कौन–सी ग्रीनहाउस गैस नहीं है ? (A) कार्बन डाइऑक्साइड (B) ओज़ोन (C) मीथेन (D) नाइट्रोजन ऑक्साइड	D	Incorrect Key	Ambiguous chemical nomenclature in option (d)	Invalid	Carbon dioxide, ozone and methane are recognized greenhouse gases. The option “nitrogen oxide” in the question refers to common nitrogen oxides such as nitric oxide/nitrogen monoxide (NO) and nitrogen dioxide (NO_2), which are generally treated as air pollutants and not listed as principal greenhouse gases at this level. Nitrous oxide, N_2O, is a greenhouse gas; however, it is specifically called nitrous oxide, not simply “nitrogen oxide.” Since the question uses the term “nitrogen oxide” and not “nitrous oxide,” the intended answer is clear in the given context. • **Greenhouse gases(GHGs)** are a group of compounds that are responsible for **absorbing heat (longwave radiation) in the atmosphere, keeping the Earth's surface warmer by allowing direct sunlight (relative short wave energy) to reach the Earth's surface.** • **Most Prominent Green House Gases**: ◦ Water vapour ◦ Carbon Dioxide, CO2 ◦ Methane, CH4 ◦ **Nitrous Oxide, N_2O (and NOT Nitrogen Oxide)** ◦ Ozone, O3 • Other GHGs are **carbon monoxide, fluorinated gases, chlorofluorocarbons (CFCs), black carbon (soot), and brown carbon.** • Amongst all the **greenhouse gases, only water vapour can absorb both incoming (UV) and outgoing (infrared) radiation.** • If GHG emissions remain unchecked, by the end of this century, the global temperatures may increase by **5^{o}C, causing global warming.**	(D)

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3	**Mathematics**	**129**	D-146	**129.** If $\vec{a} = i + j + k$ and $\vec{b} = i - j + k$, then $\vec{a} \times \vec{b}$ is _____. यदि $\vec{a} = i + j + k$ और $\vec{b} = i - j + k$ है, तो $\vec{a} \times \vec{b}$ _____ है । (A) 2 (B) $\sqrt{6}$ (C) $2\sqrt{2}$ (D) 4	C	Incorrect Question	Objection: While the question asks for the cross product of →a & →b (which would give a vector), the options given are all scalars.	Invalid	Although the question states a × b, the answer choices are all scalar quantities. It is evident from the options that the examination intended to ask for the magnitude of the cross product, which is a common convention in objective-type examinations. Given: a = i + j + k = (1, 1, 1) b = i − j + k = (1, −1, 1) a × b = (2, 0, −2) Therefore, $\|a \times b\| = \sqrt{(2^2 + 0^2 + (-2)^2)}$ $= \sqrt{8}$ $= 2\sqrt{2}$ Hence, Option (C) is the correct answer. The intent of the question is clear from the answer options, and there is no ambiguity in determining the correct response. Therefore, the objection is not sustainable.	(C)